Roger Adams and H. W. Stewart. Restricted Rotation in Arylamines. II. Preparation and Resolution of N-Succinyl-N-ethyl-3-bromomesidine and 5-Alkoxy-4-N-succinyl-4-amino-1,3-dimethylbenzenes.

Page 2860. Col. 1, line 1 of last paragraph, for "(1 mole)" read "(0.1 mole)," and same below for diethyl sulfate and sodium nitrate.

T. Q. Chou and T. T. Chu. The Preparation and Properties of Peimine and Peiminine.

Page 2936. Col. 1, line 5, for "wereas signed" read "were assigned."

George Wash, Billie Shive and H. L. Lochte. Normal and Abnormal Alkylation of 2-Methylcyclopentyl Methyl Ketone.

Page 2975. In the title, for "Methyl" read "Phenyl."—H. L. LOCHTE.

W. D. Larson and W. J. Tomsicek. The Activity Coefficients of the Undissociated Part of Weak Acids. II. Oxalic Acid.

Page 3330. Equation (4) should read

$$E^0 = E + 0.02957 \log_{10} K_1 K_2 C \left(1 - \frac{-K_1 + \sqrt{K_1^2 + 4K_1 C}}{2C}\right)$$

—W. D. LARSON.

1942, VOL. 64

W. R. Forsythe and W. F. Giauque. The Entropies of Nitric Acid and its Mono- and Tri-hydrates. Their Heat Capacities from 15 to 300°K. The Heats of Dilution at 298.1°K. The Internal Rotation and Free Energy of Nitric Acid Gas. The Partial Pressures over its Aqueous Solutions.

Pages 60 and 61. The authors call attention to the following arithmetical errors:

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (1)$$

$\Delta F^0_{298.1}$ should be −19030 cal. instead of −11,539 cal.

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (g)$$

$\Delta F^0_{298.1}$ should be −17948 cal. instead of −10.457

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (a = 1)$$

$\Delta F_{298.1}$ should be −26345 cal. instead of −18854 cal.

The above errors are repeated in the summary.
Also in the summary

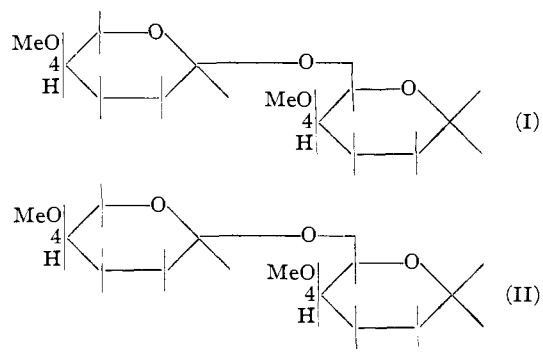
$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (g)$$

$\Delta H_{298.1}$ should be −31994 cal. instead of −3199 cal.

—W. R. FORSYTHE and W. F. GIAUQUE.

E. V. White. The Constitution of Arabo-galactan. II. The Isolation of Heptamethyl- and Octamethyl-6-galactosidogalactose through Partial Hydrolysis of Methylated Arabo-galactan.

Page 303. The author writes "Formulas 1 and 2 representing octamethyl-6-*d*-galactosidogalactose and heptamethyl-6-*d*-galactosidogalactose are incorrect in the relative positions of the methoxyl and hydrogen groups at position 4 in all the monosaccaride units, which should be reversed as indicated



Page 304. In col. 1 of the Experimental, line 30, in the formula, for "Hg" read "H_8."—E. V. WHITE.

Emmett R. Barnum and Cliff S. Hamilton. 5-Amino- and 1-Aminobenzo(f)-quinolines and Derivatives.

Page 542. Col. 2, lines 11 and 20, for "*p*-dimethylaminobenzaldehyde" read "*p*-diethylaminobenzaldehyde."—C. S. HAMILTON.

H. J. Lucas and Clark W. Gould, Jr. Brucine as a Reagent for Partially Resolving Bromoalkanes; the Configurations of Some Diastereomeric Dibromoalkanes.

Page 602. In footnote (7) for "(+1.0016)" read "(0.0016)," for "compound" read "compared" and in the citation, for "1471" read "147."

N. Howell Furman and Clark E. Bricker. A Polarographic Study of *o*-Phthalic Acid and Phthalates.

Page 665. In the legend for Fig. 6, the last four lines should read "Curve a at *p*H 4.02; curve b at *p*H 3.94; curve c at *p*H 3.75; curve d at *p*H 3.57; curve e at *p*H 3.48; curve f at *p*H 3.25; curve g at *p*H 3.06; curve h at *p*H 2.60; polarogram was made with capillary no. 3."—N. H. FURMAN.

S. C. Schumann, J. G. Aston and Malcolm Sagenkahn. The Heat Capacity and Entropy, Heats of Fusion and Vaporization and the Vapor Pressures of Isopentane.

Page 1041. The authors write "In Table IV the molal heat of vaporization of isopentane at 298.16°K. calculated from Eq. (1) and the modified Berthelot equation at 298.16°K. with T_c = 461°K. and P_c = 32.9 atm. is 5965 calories (Berthelot correction = 239 calories), instead of the value given."

Page 1043. "In Table VII the free energy of neopentane was accidentally computed using the free rota-

tional entropy. Using the correct value for this quantity [Aston and Messerly, THIS JOURNAL, **58**, 2354 (1936)] Table VII should then be:

FREE ENERGY OF FORMATION OF THE PENTANES AT 298.16°K.

	ΔF_{298}, calories	% in equilibrium mixture	% if no neopentane
Neopentane	−3289 ± 320	48.7	...
Isopentane	−3243 ± 300	48.3	93.7
n-Pentane	−1647 ± 270	3.0	6.3

These changes do not affect the rest of the paper, since no other numerical values or conclusions were derived from the erroneous values.—S. C. SCHUMANN, J. G. ASTON AND M. L. SAGENKAHN.

Joseph R. Stevens, Ralph H. Beutel and Earl Chamberlin. 3,4-Substituted Pyridines. I. Synthesis of 3-Vinyl-4-methylpyridine.

Page 1093. The absorption spectrum in Fig. 1 is for Compound VI instead of V, and in Fig. 2 for Compound VII instead of VI.—R. T. MAJOR.

H. A. Laitinen. The Potential of the Ytterbic-Ytterbous Ion Electrode.

Page 1135. Col. 2, line 13, for "−1.69 volts" read "−1.169 volts."—H. A. LAITINEN.

Arthur C. Cope and Evelyn M. Hancock. Synthesis of 2-Alkylaminoethanols from Ethanolamine.

Pages 1504 (Table I) and 1505 (line 7 of the Experimental Part), for "2,2,6-" read "3,3,5-trimethylcyclohexyl" and "3,3,5-trimethylcyclohexanone."—ARTHUR C. COPE.

James W. McBain and A. M. Soldate. The Solubility of Propylene Vapor in Water as Affected by Typical Detergents.

Page 1556. Heading of Table I, for "10^{-5}" read "10^{-7}." O'Connor (ref. 1) obtained 3.8, 3.8, 3.7 and 3.5 × 10^{-7} g. of propylene per gram of water per millimeter pressure. —J. W. MCBAIN.

Bradford P. Geyer with George McP. Smith. Preparation and Properties of Some Peri-hydroxyquinone Inner Complexes.

Page 1649. In col. 2, line 2, for "chloroform" read "chlorobenzene."—B. P. GEYER and G. MCP. SMITH.

A. Polgár and L. Zechmeister. Isomerization of β-Carotene. Isolation of a Stereoisomer with Increased Adsorption Affinity.

Page 1858. Line 23, omit the word "no."—L. ZECHMEISTER.

W. D. Kumler and George M. Fohlen. The Dipole Moment and Structure of Urea and Thiourea.

Page 1945. In Table II for "*unsym*-Diphenylure" read "*unsym*-Diphenylurea."

Page 1946. Both formulas at the end of col. 1 should have a single bond between carbon and oxygen.

Page 1947. In the middle of col. 2, the last structure of thiourea should have the charge removed over the nitrogen that is singly bonded to carbon.—W. D. KUMLER.

R. P. Linstead and W. E. Doering. The Stereochemistry of Catalytic Hydrogenation. II. The Preparation of the Six Inactive Perhydrodiphenic Acids.

Page 1993. In Col. 2, first line of the diagram, for "Diphenyl" read "Dimethyl."—R. P. LINSTEAD.

Page 1994. In formula II, there should be a third black dot in the blank "hole" in the formula.

R. P. Linstead and W. E. Doering. The Stereochemistry of Catalytic Hydrogenation. III. Optically Active Perhydrodiphenic Acids. A Proof of the Configuration of the Backbone.

Page 2004. Col. 1, in the second formula there should be another black dot in the space in the right-hand ring.

R. P. Linstead and Selby B. Davis. The Stereochemistry of Catalytic Hydrogenation. IV. Hexahydrodiphenic Acids.

Page 2007. In formula II a black dot is missing in the right-hand ring of formula II.

R. P. Linstead, Richard R. Whetstone and Philip Levine. The Stereochemistry of Catalytic Hydrogenation. VI. The Hydrogenation of 9-Phenanthrol and Related Substances and the Identification of Three of the Possible Stereoisomeric Forms of the Perhydrophenanthrene Ring.

Page 2017. Col. 1 in, the second line above the table, insert the word "table" between "The" and "below." Also, in Col. 2, in the table under "Alcohols," the first compound should be "*sym*-octahydro-9-phenanthrol."—R. P. LINSTEAD.

M. L. Wolfrom and P. W. Morgan. O-Pentaacetyl-*d*-gluconates of Polyhydric Alcohols and Cellulose.

Page 2026. Column 1, lines 9 and 16, for "*dextro*-sorbitol" read "*levo*-sorbitol," as the ordinary form is meant.

Page 2027. In Table I the fifth entry should read *levo* instead of *dextro*, and the same change should be made in line 11 of col. 1.

Page 2028. In the first line of paragraph 2 of the Summary read *levo* for *dextro*.—M. L. WOLFROM.

Alfred Saffer and T. W. Davis. Products from the Wurtz Reaction and the Mechanism of their Formation.

Page 2039. The apparatus diagram which should have been included in the experimental section was inadvertently omitted, and is printed herewith.